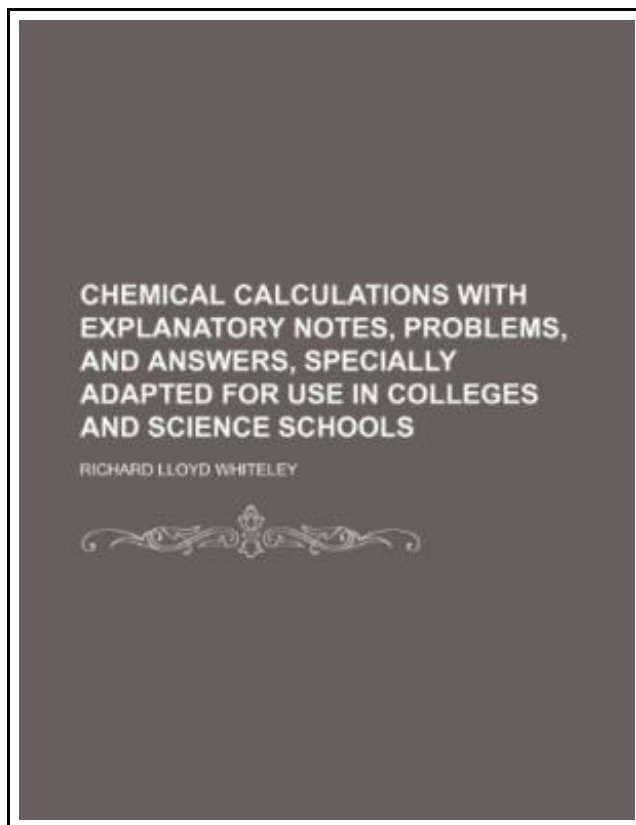


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Rarebookclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1896 edition. Excerpt: .are required. Calculate the percentage of chlorine in the specimen. The equations representing the reactions are-- $KClO_3 + 3H_2 = KC_1 + 3H_2O$ ;  $KCl + AgNO_3 = AgCl + KNO_3$ . Since the mol. wgt. of  $AgNO_3 = 170$ , a normal solution will contain 170 grams in 1 litre of solution; 1 c.c. will contain 0.170 gram  $AgNO_3$ . If the solution is deci-normal, since only--of 170 J 10 grams is dissolved, 1 c.c. will equal 0.0170 gram  $AgNO_3$ . Now 170 grams  $AgNO_3$  will precipitate 35.5 grams of chlorine as  $AgCl$ ;-. 1 c.c. of  $AgNO_3 (= 0.17 \text{ gram})$  will precipitate (or is equivalent to) 0.0355 gram of chlorine; and. . 1 c.c. of  $AgNO_3 (= 0.017 \text{ gram})$  will precipitate (or is equivalent to) 0.00355 gram of chlorine. N In the example given, 26 c.c. of  $AgNO_3$  were used;  $26 \times 0.00355 =$  the weight of Cl, to which they are equivalent; and, since that is obtained from 0.32 gram  $KClO_3$ , the percentage amount is found by the following proportion: --0.32: 100::  $26 \times 0.00355$ : x;  $100 \times 26 \times 0.00355$  oa a. x =-----= 26.64: per cent, chlorine. L)-o-Z-Example 6.--100 c.c. of a solution of sulphurous acid re-N quired 5-15 c.c. of--iodine solution (1 c.c. = 0.0127 gm.) to completely oxidise it. Find the weight of  $SfX$  per litre of solution.  $SO_2 + I_2 + 2H_2O = H_2SO_4 + 2HI$ . 1 To titrate a solution means to ascertain the titre, or strength, of it by volumetric methods. Since one molecule of iodine...



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